

Ionic Bonds and Compounds Worksheet #1

Fill in the oxidation number in the space provided next to each element. Then write the chemical formula for each compound; Na_2O is done as an example.

		Non-Metals			
		O <u>-2</u>	F —	Cl —	Se —
Metals	Na <u>+1</u>	Na_2O			
	Ba —				
	Ca —				
	Rb —				
	In —				

Name the following ionic compounds:

NaI _____

CaCl_2 _____

Li_2O _____

AlF_3 _____

KI _____

Mg_3N_2 _____

CaO _____

AlN _____

Over →

Give the chemical formula for the following ionic compounds. Remember to use the correct subscripts!

Beryllium sulfide: _____

Aluminum bromide: _____

Calcium chloride: _____

Strontium nitride: _____

Potassium oxide: _____

Sodium nitride: _____

Level 1:

Iron(III) oxide: _____

Iron(II) oxide: _____

Copper(I) fluoride: _____

Copper(II) fluoride: _____

Name the following ionic compounds. Remember to use the correct Stock Number (roman numeral) for the transition metals.

MnF₃ _____

MnF₂ _____

Cu₂O₃ _____

CrO _____

CoCl₂ _____

Cr₂S₃ _____

FeCl₃ _____

CoO _____